Morning Assignment 03 jun 2026
Chemical Kinetics Module-kin.03

Fill in the blanks

1. The expression relating reaction rate with concentration of reactants is called the _____ law.
2. In the rate equation, k represents the rate _____.
3. The sum of powers of concentration terms in a rate equation is called the _____ of reaction.
4. A reaction whose rate is independent of concentration is called a _____ order reaction.
5. For the reaction Rate = $k[A]^2[B]$, the overall order is _____.
6. The rate law of a reaction is determined _____.
7. Reactions occurring in a single step are called _____ reactions.
8. Reactions occurring through several steps are called _____ reactions.
9. The unit of rate constant for a first order reaction is _____.
10. The unit of rate constant for a second order reaction is _____.

Short Question Answers

1. What is a rate law?
2. What is meant by rate constant?
3. What is the order of a reaction?
4. What is a zero order reaction?
5. Why cannot rate law be predicted from a balanced chemical equation?

6. What is an elementary reaction?
7. What is a complex reaction?
8. What is the unit of rate constant for a first order reaction?
9. What is the unit of rate constant for a second order reaction?
10. What does the exponent of concentration terms indicate in a rate equation?

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